**3.3 – Net Ionic Equations – Practice**

Try the following questions. Write the balanced equation (CRUCIAL before continuing), total ionic equation, and net ionic equation.

**1.  AgNO3 (aq)  +  KCl (aq)   🡪    AgCl (s)  +  KNO3 (aq)**

**2.  Mg(NO3)2 (aq)   +   Na2CO3 (aq)   🡪   MgCO3 (s)  +   NaNO3 (aq)**

**3.  strontium bromide (aq)  +  potassium sulfate (aq)  🡪  strontium sulfate (s)  +  potassium bromide (aq)**

**4.  manganese (II) chloride (aq)  +   ammonium carbonate (aq)  🡪  manganese (II) carbonate (s)  +  ammonium chloride (aq)**

**5.  chromium (III) nitrate (aq)  +  iron(II) sulfate (aq)  🡪  chromium(III) sulfate (aq) + iron(II) nitrate (aq)**

**6.  K3PO4 (aq) +  Al(NO3)3 (aq)  🡪  AlPO4 (s)  +  KNO3 (aq)**

**7.  BeI2 (aq) + Cu2SO4 (aq) 🡪  BeSO4 (aq)  +  CuI (s)**

**8.  Ni(NO3)3 (aq)  +  KBr (aq)  🡪  NiBr3 (aq)  +  KNO3 (aq)**

**9.  cobalt (III) bromide (aq)  +  potassium sulphide (aq)  🡪 cobalt (III) sulphide (s) + potassium bromide (aq)**

**10.  barium nitrate (aq)   +   ammonium phosphate (aq)  🡪 barium phosphate (s) + ammonium nitrate (aq)**

**11. calcium hydroxide  +  iron (III) chloride 🡪 calcium chloride (aq) + iron (III) hydroxide (s)**

**12.  rubidium fluoride  +  copper (II) sulfate  🡪 rubidium sulphate (aq) + copper (II) fluoride (aq)**